

A solid mixture weighing 0.5485g contained only ferrous ammonium sulfate

Question: A solid mixture weighing 0.5485 g contained only ferrous ammonium sulfate hexahydrate and ferrous chloride hexahydrate. The sample was dissolved in 1M H₂SO₄, ...

A solid mixture weighing 0.05485 g contained only ferrous ammonium sulfate and ferrous chloride. The sample was dissolved in 1 M H₂SO₄, and the Fe²⁺ required 13.39 mL of ...

A solid mixture weighing 0.05485 g contained only ferrous ammonium sulfate and ferrous chloride. The sample was dissolved in 1 M H₂SO₄, and the Fe²⁺ required 13.39 mL of 0.01234 M Ce⁴⁺ for complete oxidation to Fe³⁺ (Ce⁴⁺ + Fe²⁺ ...

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Make sure you show complete procedures a) A solid mixture weighing 0.05485 g contained only ferrous ammonium sulfate and ferrous chloride. The sample was dissolved in 1 M H₂SO₄, and the Fe²⁺ required 13.39 mL of 0.01234 M Ce⁴⁺ ...

Solution For A solid mixture weighing 0.05485 g contained only ferrous ammonium sulphate and ferrous chloride. The sample was dissolved in 1M H₂SO₄, and the Fe²⁺ required 13.39 mL of ...

Here we have a solid mixture which weighs 0.5485 g and contains only ferrous ammonium sulfate hexahydrate and ferrous chloride hexahydrate. We need to calculate the weight percent of Cl ...

A solid mixture weighing 0.05485 g contained only ferrous ammonium sulfate hexahydrate and ferrous chloride hexahydrate. The sample was dissolved in 1.0 M H₂SO₄, and the Fe²⁺ required 13.39 mL of 0.01234 M Ce⁴⁺ for complete ...

In our exercise, stoichiometry is used to calculate the moles of (Fe^{3+}), which comes from the complete oxidation of (Fe^{2+}) in ferrous ammonium sulfate and ferrous chloride. ...

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VIDEO ANSWER: So for this problem, we have a series of chemical reactions, which involves some iron. And, we're told, is that we have some potassium chlorate but then decomposes. It's ...

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A solid mixture weighing 0.515 g contained only ferrous ammonium sulfate and ferrous chloride. The same was dissolved in 2 M H₂SO₄, and the Fe²⁺ required 12.85 mL ...

A solid mixture weighing 0.5485 g contained only ferrous ammonium sulfate hexahydrate (FM 392.13) and ferrous chloride hexahydrate (FM 234.84). The sample was dissolved in 1M ...

A solid mixture weighing 0.5485 g contained only ferrous ammonium sulfate hexahydrate and ferrous chloride hexahydrate. The sample was dissolved in 1M H₂SO₄, oxidized to Fe³⁺ ...

A solid mixture weighing 0.05485 g contained only ferrous ammonium sulphate and ferrous chloride. The sample was dissolved in 1 M H₂SO₄, and the Fe²⁺ required 13.39 mL ...

A solid mixture weighing 0.5485 g contained only ferrous ammonium sulfate hexahydrate and ferrous chloride hexahydrate. The sample was dissolved in 1 M ...

Web: <https://bardzyndzalek.olsztyn.pl>

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